

Solubility-Product Constants for Compounds at 25°C

Name	Formula	K _{sp}	Name	Formula	K _{sp}
Barium carbonate	BaCO ₃	5.0 x 10 ⁻⁹	Lead (II) chromate	PbCrO ₄	2.8 x 10 ⁻¹³
Barium chromate	BaCrO ₄	2.1 x 10 ⁻¹⁰	Lead (II) fluoride	PbF ₂	3.6 x 10 ⁻⁸
Barium fluoride	BaF ₂	1.7 x 10 ⁻⁶	Lead (II) sulfate	PbSO ₄	6.3 x 10 ⁻⁷
Barium oxalate	BaC ₂ O ₄	1.6 x 10 ⁻⁶	Lead (II) sulfide*	PbS	3.0 x 10 ⁻²⁸
Barium phosphate	Ba ₃ (PO ₄) ₂	3.4 x 10 ⁻²³	Magnesium carbonate	MgCO ₃	3.5 x 10 ⁻⁸
Barium sulfate	BaSO ₄	1.1 x 10 ⁻¹⁰	Magnesium hydroxide	Mg(OH) ₂	1.6 x 10 ⁻¹²
Cadmium carbonate	CdCO ₃	1.8 x 10 ⁻¹⁴	Magnesium oxalate	MgC ₂ O ₄	8.6 x 10 ⁻⁵
Cadmium hydroxide	Cd(OH) ₂	2.5 x 10 ⁻¹⁴	Manganese (II) carbonate	MnCO ₃	5.0 x 10 ⁻¹⁰
Cadmium sulfide*	CdS	8.0 x 10 ⁻²⁸	Manganese (II) hydroxide	Mn(OH) ₂	1.6 x 10 ⁻¹³
Calcium carbonate	CaCO ₃	4.5 x 10 ⁻⁹	Manganese (II) sulfide*	MnS	2.0 x 10 ⁻⁵³
Calcium chromate	CaCrO ₄	7.1 x 10 ⁻⁴	Mercury (I) chloride	Hg ₂ Cl ₂	1.2 x 10 ⁻¹⁸
Calcium fluoride	CaF ₂	3.9 x 10 ⁻¹¹	Mercury (I) iodide	Hg ₂ I ₂	1.1 x 10 ⁻²⁸
Calcium hydroxide	Ca(OH) ₂	6.5 x 10 ⁻⁶	Mercury (II) sulfide*	HgS	2.0 x 10 ⁻⁵³
Calcium phosphate	Ca ₃ (PO ₄) ₂	2.0 x 10 ⁻²⁹	Nickel (II) carbonate	NiCO ₃	1.3 x 10 ⁻⁷
Calcium sulfate	CaSO ₄	2.4 x 10 ⁻⁵	Nickel (II) hydroxide	Ni(OH) ₂	6.0 x 10 ⁻¹⁶
Chromium (III) hydroxide	Cr(OH) ₃	1.6 x 10 ⁻³⁰	Nickel (II) sulfide*	NiS	3.0 x 10 ⁻²⁰
Cobalt (II) carbonate	CoCO ₃	1.0 x 10 ⁻¹⁰	Silver bromate	AgBrO ₃	5.5 x 10 ⁻⁵
Cobalt (II) hydroxide	Co(OH) ₂	1.3 x 10 ⁻¹⁵	Silver bromide	AgBr	5.0 x 10 ⁻¹³
Cobalt (II) sulfide*	CoS	5.0 x 10 ⁻²²	Silver carbonate	Ag ₂ CO ₃	8.1 x 10 ⁻¹²
Copper (I) bromide	CuBr	5.3 x 10 ⁻⁹	Silver chloride	AgCl	1.8 x 10 ⁻¹⁰
Copper (II) carbonate	CuCO ₃	2.3 x 10 ⁻¹⁰	Silver chromate	Ag ₂ CrO ₄	1.2 x 10 ⁻¹²
Copper (II) hydroxide	Cu(OH) ₂	4.8 x 10 ⁻²⁰	Silver iodide	AgI	8.3 x 10 ⁻¹⁷
Copper (II) sulfide*	CuS	6.0 x 10 ⁻³⁷	Silver sulfate	Ag ₂ SO ₄	1.5 x 10 ⁻⁵
Iron (II) carbonate	FeCO ₃	3.2 x 10 ⁻¹¹	Silver sulfide*	Ag ₂ S	6.0 x 10 ⁻⁵¹
Iron (II) hydroxide	Fe(OH) ₂	7.9 x 10 ⁻¹⁶	Strontium carbonate	SrCO ₃	9.3 x 10 ⁻¹⁰
Lanthanum fluoride	LaF ₃	2.0 x 10 ⁻¹⁹	Tin (II) sulfide*	SnS	1.0 x 10 ⁻²⁶
Lanthanum iodate	La(IO ₃) ₃	6.1 x 10 ⁻¹²	Zinc carbonate	ZnCO ₃	1.0 x 10 ⁻¹⁰
Lead (II) carbonate	PbCO ₃	7.4 x 10 ⁻¹⁴	Zinc hydroxide	Zn(OH) ₂	3.0 x 10 ⁻¹⁶
Lead (II) chloride	PbCl ₂	1.6 x 10 ⁻⁵	Zinc oxalate	ZnC ₂ O ₄	2.7 x 10 ⁻⁸
			Zinc sulfide*	ZnS	2.0 x 10 ⁻²⁵

*Equilibrium reaction: $metal\ sulfide_{(s)} \leftrightarrow metal\ ion_{(aq)} + HS_{(aq)} + OH_{(aq)}$